

CHAPTER 18 NOTES
DEFINING ACIDS AND BASES
p.595 – p.623

Section 18-1

	Acids	Bases
Taste		
Touch		
Reaction with metals		
Electrical conductivity		
Indicator (litmus paper)		

What is a neutralization reaction?

What forms from a neutralization reaction?

What is the Arrhenius Definition of Acids and Bases?

What is produced when an acid and a metal react?

State 3 reasons why the Arrhenius definition of acids and bases is limited.

What is the Bronsted-Lowry definition of acids and bases?

What is a proton donor and proton acceptor?

Why is sulfuric acid (H_2SO_4) described as a diprotic acid?

What is a hydronium ion?

Define conjugate acids and conjugate bases and provide examples of each.

Section 18-2

What is considered a strong acid?

Which acid would have a higher conductivity, a strong one or weak one? Explain.

What is considered a strong base?

Which kind of base – strong or weak- reacts to completion with water?

If a strong base is added to a strong acid, what will happen?

What does the acid and base dissociation constant indicate?

Do practice problems #1 and #2 on page 613.

What is a salt hydrolysis reaction?

Identify the four kinds of salts and state whether they form acidic, basic, or neutral solutions in water.

Section 18-3

What is an acidic hydrogen?

Describe a structural feature that is common to all bases.

State the 3 rules to naming acids.